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Give a complete answer. Chem-271/Precipitation
Reactions Lab/Page 3 (10/08) Solubility Rules for Salts
in Water 1. All lithium, sodium, potassium and
ammonium salts are soluble regardless of the counter
ion. 2. All nitrates are soluble regardless of the

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counter ion. 3. All chlorides, bromides and iodides are soluble, except when the counter ion is

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Post-Lab Questions. Write the balanced chemical reaction for reaction 1. Be sure to include the states of matter and correct formulas. If applicable (ie: if you classified it as a precipitation reaction) include the net ionic equation as well. Write the balanced chemical reaction for reaction 3.

Precipitation Reactions Worksheet Key

$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$ Observing precipitation reactions can be useful in the laboratory to determine the presence of various ions in solution. For instance, if silver nitrate is added to a solution of an unknown salt and a precipitate is observed, the unknown solution might contain chloride (Cl^-).

4.06 lab - Salts Precipitation Reactions Lab Report Lesson ...

Salts: Precipitation Reactions Lab Report Lesson 4.06
I. Objective: you will mix aqueous solutions of differing chemical natures to observe which do or do not form a precipitate. A color change or the formation of a precipitate is a clear signal that a chemical reaction has occurred and a salt has formed.
II.

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Precipitation Reactions Worksheet Key For each of the following reactants, predict whether a precipitation reaction will take place between them. If there is no reaction, write "no reaction". If there is a reaction, write the complete, complete ionic, and net ionic equations that describe the reaction. a. $\text{Li}_2\text{CO}_3(\text{aq}) + \text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2(\text{aq}) \dots$

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The result is a solid precipitate that rapidly comes out of solution. For example, when solutions of silver nitrate, AgNO_3 , and sodium chloride, NaCl are combined, a double displacement reaction...

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Answer Key Precipitation Lab. advertisement. Answer Key Precipitation Lab Purpose: 1. Identify the ions present in various aqueous solutions. 2. Combine solutions and identify the reactions that form precipitates. 3. Write balanced molecular equation, complete ionic equation, and a net ionic equation for each precipitation reaction. Answer Key ...

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Semester 1 Unit 4 LAB...

A precipitation reaction results in the formation of an insoluble product. Whether a precipitate, an insoluble solid that separates from the solution, will form depends on the solubility of the solute. Precipitation reactions usually involve ionic compounds, and although all ionic compounds are strong electrolytes they are not equally soluble.

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as the result of a chemical reaction, the observer will see a precipitate; therefore she can predict many reactions simply by noticing the possible new combinations of ions. Also, a knowledge of common weak electrolytes allows you to predict reactions that make new molecular products, like H_2O , or unstable compounds like $\text{H}_2\text{CO}_3(\text{aq})$.

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